



SRI BHARATHI

ENGINEERING COLLEGE FOR WOMEN

(Approved by AICTE, New Delhi and Affiliated to Anna University, Chennai)
Kaikkurichi, Pudukkottai -622 303

www.sbec.edu.in

NAAC DOCUMENTS



Quality Indicator Frame Work

Criterion – 1

CURRICULAR ASPECTS

Submitted by

IQAC

Internal Quality Assurance Cell

Sri Bharathi Engineering College for Women



SRI BHARATHI ENGINEERING COLLEGE FOR WOMEN

(Approved by AICTE, New Delhi, Affiliated to Anna University, Chennai-25)

Kaikkurichi, Pudukkottai, Tamil Nadu – 622 303, India

Criterion 1	Curricular Aspects	100
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1.1 Curricular Planning and Implementation(20)

1.1.1 The Institution ensures effective curriculum planning and delivery through a well-planned and documented process including Academic calendar and conduct of continuous internal Assessment

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Kaikkurichi, Pudukkottai, Tamil Nadu – 622 303, India

DEPARTMENT OF SCIENCE AND HUMANITIES

PREFACE OF THE COURSE FILE

Batch : 2018-2022

Academic Year : 2018-2019 / ODD

Program : COMPUTER SCIENCE AND ENGINEERING

Year & Semester : 1st Year / 1st Semester / 'A' Section

Course Code : CY8151 NBA Course Code: C104

Name of the Course : ENGINEERING CHEMISTRY

Faculty in-charge : Mrs.T.ANNALAKSHMI

Signature of the Faculty Incharge

Dr. S.THILAGAVATHI M.E.,Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
Kaikkurichi - 622 303, Pudukkottai Dt.

HOD / S&H
HOD / S&H

SRI BHARATHI ENGINEERING
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PUDUKKOTTAI - 622 303.

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DEPARTMENT OF SCIENCE AND HUMANITIES

REVIEW OF COURSE FILE

(To be pasted on the inner side of the file-backside).(#-State Yes/No.)

S.No	Details Date:	R-I-*	R-II- *&	R-III- *&	R-IV- *&\$	R-V- *&\$@
1.	Preface of the course file	yes				
2.	Vision, Mission, PEOs, POs, PSOs, Blooms taxonomy	yes				
3.	Subject handlers of yesteryears					
4.	Timetable/Workload of the staff – Distribution of teaching load – Roles and Responsibilities	yes				
5.	Syllabus signed by staff & HoD	yes				
6.	Lecture Schedule signed by staff & HoD	yes				
7.	Course Committee meeting circular and minutes	yes				
8.	Identification of Curricular gap and Content Beyond the syllabus	yes				
9.	Self-study topics	yes				
10.	Previous AU Question papers	yes				
11.	Unit wise Q&A and Objective type questions	yes				
12.	Unit wise course material	yes				
13.	Assignment question paper with sample answer sheets and mark entry		yes			
14.	Tutorial question paper with key and mark entry		yes			
15.	Class test/IA test Q Paper with Key, sample answer papers and mark entry		yes			
16.	IA Test- result analysis-CAP-evidence-root cause analysis.		yes			
17.	Retest –Q paper-Attendance-marks		yes			
18.	AU Web portal entry sheet		yes			
19.	Very poor performance in first two tests-action taken.-communication to parents-evidence					
20.	Absence for two tests-action taken-communication to parents-evidence.					
21.	Indiscipline of student reported, if any					
22.	Special class/coaching class/remedial class/attendance-CAP					
23.	Conduct of Seminar, Quizzes - proof					
24.	Content beyond the syllabus - proof			yes		
25.	Student feedback on faculty			yes		
26.	Course end survey					
27.	Internal Assessment sheet			yes		
28.	AU question paper with students feedback					
29.	Discrepancy of the question paper and correspondence, if any					
30.	AU result analysis-Details of arrear students.					
31.	AU grade sheet					yes
32.	CO – PO & PSO attainment sheet					yes
	Signature of Course handling faculty					
	Signature of HoD/ S&H Dr. S. THILAGAVATHI M.E., Ph.D. PRINCIPAL					

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DEPARTMENT OF SCIENCE AND HUMANITIES

ACADEMIC YEAR (2018 – 2019) -ODD SEMESTER

INDIVIDUAL STAFF WORKLOAD

S. NO	STAFF NAME	SUBJECT CODE & NAME	YEAR & DEPT	NO OF HOURS	TOTAL HOURS
1	Mrs.N.Vithya	CS6702-Graph theory Applications	IV /CSE	6	18
		MA6351- Discrete Mathematics	II / CSE&IT	6	
		MA8151-Engineering Mathematics-I	I-(SEC-A)/ CSE	6	
2	Ms.S.Poonguzhali	CS6704 - Resource Management Techniques	IV /CSE	6	18
		MA6566-Discrete Mathematics	III/CSE	6	
		MA6351-Transform and PDE	II /CIVIL&EEE	6	
3	Mrs.M.Vanitha	MA8352-Linear Algebra and PDE	II / ECE	6	12
		MA8151- Engineering Mathematics-I	I-(SEC-B) CIVIL,ECE&EEE	6	
4	Mrs.R.Saratha	PH8151-Engineering Physics	I-SEC-A /CSE	05	05
5	Mrs.V.Vinojini	BS8161-Physics Laboratory	I (SEC-A&SEC-B) CIVIL,CSE,ECE&EE E	06	06
6	Mrs.T.Renugadevi	PH8151-Engineering Physics	I-SEC-B CIVIL,ECE&EEE	05	05
	Ms.T.Annalakshmi	CY8151-Engineering Chemistry	I SEC-A /CSE	05	08
		BS8161-Chemistry Laboratory	I SEC-A/CSE	03	
7	Mrs.S.Renugadevi	CY8151-Engineering Chemistry	I SEC-B/ CIVIL,ECE&EEE	05	14
		BS8161-Chemistry Laboratory	I SEC-B/ CIVIL,ECE&EEE	03	
		GE6351-Environmental Science and Engineering	III /ECE	6	
9	Mr.S.Ramesh Raja	HS8151-Communicative English	I-SEC-A/CSE	05	10
		GE2321-Communication Skill Laboratory	III/CIVIL,EEE,IT	3	
		HS8381-Interpersonal Skills/Listening &Speaking	II/CSE	2	
10	Mrs.P.Alagumathi	HS8151-Communicative English	I-SEC-B CIVIL,ECE&EEE	05	09
		HS8381-Interpersonal Skills/Listening &Speaking	II/CIVIL,ECE	04	

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UNIT I WATER AND ITS TREATMENT 9

Hardness of water – types – expression of hardness – units – estimation of hardness of water by EDTA – numerical problems – boiler troubles (scale and sludge) – treatment of boiler feed water – Internal treatment (phosphate, colloidal, sodium aluminate and calgon conditioning) external treatment – Ion exchange process, zeolite process – desalination of brackish water - Reverse Osmosis.

UNIT II SURFACE CHEMISTRY AND CATALYSIS 9

Adsorption: Types of adsorption – adsorption of gases on solids – adsorption of solute from solutions – adsorption isotherms – Freundlich's adsorption isotherm – Langmuir's adsorption isotherm – contact theory – kinetics of surface reactions, unimolecular reactions, Langmuir - applications of adsorption on pollution abatement. Catalysis: Catalyst – types of catalysis – criteria – autocatalysis – catalytic poisoning and catalytic promoters - acid base catalysis – applications (catalytic convertor) – enzyme catalysis– Michaelis – Menten equation.

UNIT III ALLOYS AND PHASE RULE 9

Alloys: Introduction- Definition- properties of alloys- significance of alloying, functions and effect of alloying elements- Nichrome and stainless steel (18/8) – heat treatment of steel. Phase rule: Introduction, definition of terms with examples, one component system -water system - reduced phase rule - thermal analysis and cooling curves - two component systems - lead-silver system- Pattinson process.

UNIT IV FUELS AND COMBUSTION 9

Fuels: Introduction - classification of fuels - coal - analysis of coal (proximate and ultimate) - carbonization - manufacture of metallurgical coke (Otto Hoffmann method) - petroleum - manufacture of synthetic petrol (Bergius process) - knocking - octane number - diesel oil - cetane number - natural gas - compressed natural gas (CNG) - liquefied petroleum gases (LPG) - power alcohol and biodiesel. Combustion of fuels: Introduction - calorific value - higher and lower calorific values- theoretical calculation of calorific value - ignition temperature - spontaneous ignition temperature - explosive range - flue gas analysis (ORSAT Method).

UNIT V ENERGY SOURCES AND STORAGE DEVICES 9

Nuclear fission - controlled nuclear fission - nuclear fusion - differences between nuclear fission and fusion - nuclear chain reactions - nuclear energy - light water nuclear power plant - breeder reactor - solar energy conversion - solar cells - wind energy. Batteries, fuel cells and supercapacitors: Types of batteries – primary battery (dry cell) secondary battery (lead acid battery, lithium-ion-battery) fuel cells – H₂-O₂ fuel cell.

TOTAL: 45 PERIODS

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E.mail:sribharathienggcollege@gmail.com Website: www.sbec.edu.in

DEPARTMENT OF SCIENCE AND HUMANITIES

COURSE PLAN

Subject code	: CY8151	Year / Semester/Branch	: I/ I/CSE
Subject Name	: ENGINEERING CHEMISTRY	Batch	: 2018-2022
Staff Name	: T.ANNALAKSHMI	Academic year	: 2018-2019

COURSE OBJECTIVES:

- To make the students conversant with boiler feed water requirements, related problems and water treatment techniques.
- To develop an understanding of the basic concepts of phase rule and its applications to single and two component systems and appreciate the purpose and significance of alloys.
- Preparation, properties and applications of engineering materials.
- Types of fuels, calorific value calculations, manufacture of solid, liquid and gaseous fuels.
- Principles and generation of energy in batteries, nuclear reactors, solar cells, wind mills and fuel cells.

TEXT BOOKS:

1. S. S. Dara and S. S. Umare, "A Textbook of Engineering Chemistry", S. Chand & Company LTD, New Delhi, 2015
2. P. C. Jain and Monika Jain, "Engineering Chemistry" Dhanpat Rai Publishing Company (P) LTD, New Delhi, 2015
3. S. Vairam, P. Kalyani and Suba Ramesh, "Engineering Chemistry", Wiley India PVT, LTD, New Delhi, 2013.

REFERENCES:

1. Friedrich Emich, "Engineering Chemistry", Scientific International PVT, LTD, New Delhi, 2014.
2. Prasanta Rath, "Engineering Chemistry", Cengage Learning India PVT, LTD, Delhi, 2015.
3. Shikha Agarwal, "Engineering Chemistry-Fundamentals and Applications", Cambridge University Press, Delhi, 2015

WEB ADDRESS

https://link.springer.com/chapter/10.1007/978-94-011-7352-0_7

<https://www.publichealthnotes.com/importance-of-water-purification-and-different-methods/>

<https://www.electronicsforu.com/technology-trends/learn-electronics/different-types-of-batteries>

https://en.wikipedia.org/wiki/Solar_energy_conversion

<https://www.encyclopedie-energie.org/en/desalination-sea-water-brackish-water/>

- BB BLACK BOARD
- PPT POWER POINT PRESENTATION

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PRINCIPAL

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CY3151 ENGINEERING CHEMISTRY

LTPC
3003

UNIT I WATER AND ITS TREATMENT

9

Hardness of water – types – expression of hardness – units – estimation of hardness of water by EDTA – numerical problems – boiler troubles (scale and sludge) – treatment of boiler feed water – Internal treatment (phosphate, colloidal, sodium aluminate and calgon conditioning) external treatment – Ion exchange process, zeolite process – desalination of brackish water - Reverse Osmosis.

UNIT II SURFACE CHEMISTRY AND CATALYSIS

9

Adsorption: Types of adsorption – adsorption of gases on solids – adsorption of solute from solutions – adsorption isotherms – Freundlich's adsorption isotherm – Langmuir's adsorption isotherm – contact theory – kinetics of surface reactions, unimolecular reactions, Langmuir - applications of adsorption on pollution abatement. Catalysis: Catalyst – types of catalysis – criteria – autocatalysis – catalytic poisoning and catalytic promoters - acid base catalysis – applications (catalytic convertor) – enzyme catalysis– Michaelis – Menten equation.

UNIT III ALLOYS AND PHASE RULE

9

Alloys: Introduction- Definition- properties of alloys- significance of alloying, functions and effect of alloying elements- Nichrome and stainless steel (18/8) – heat treatment of steel. Phase rule: Introduction, definition of terms with examples, one component system -water system - reduced phase rule - thermal analysis and cooling curves - two component systems - lead-silver system- Pattinson process.

UNIT IV FUELS AND COMBUSTION

9

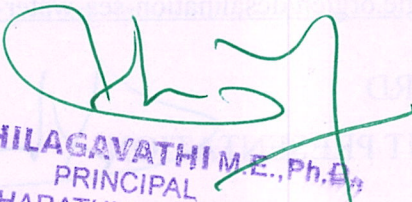
Fuels: Introduction - classification of fuels - coal - analysis of coal (proximate and ultimate) - carbonization - manufacture of metallurgical coke (Otto Hoffmann method) - petroleum - manufacture of synthetic petrol (Bergius process) - knocking - octane number - diesel oil - cetane number - natural gas - compressed natural gas (CNG) - liquefied petroleum gases (LPG) - power alcohol and biodiesel. Combustion of fuels: Introduction - calorific value - higher and lower calorific values- theoretical calculation of calorific value - ignition temperature - spontaneous ignition temperature - explosive range - flue gas analysis (ORSAT Method).

UNIT V ENERGY SOURCES AND STORAGE DEVICES

9

Nuclear fission - controlled nuclear fission - nuclear fusion - differences between nuclear fission and fusion - nuclear chain reactions - nuclear energy - light water nuclear power plant - breeder reactor - solar energy conversion - solar cells - wind energy. Batteries, fuel cells and supercapacitors: Types of batteries – primary battery (dry cell) secondary battery (lead acid battery, lithium-ion-battery) fuel cells – H₂-O₂ fuel cell.

TOTAL: 45 PERIODS


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Topic No	Topic Name	T/R*	Page No	Teaching Methodology	No of periods required	Cumulative periods
		Book				
UNIT-I WATER AND ITS TREATMENT (9)						
1.	Hardness of water , types Importance of water and its purification methods (CBS)	T1 W2	1.1-1.2	BB PPT	1	1
2.	expression of hardness ,units , estimation of hardness of water by EDTA	T1	1.2-1.6	BB	1	2
3.	numerical problems , boiler troubles (scale and sludge)	T1	1.6-1.11	BB	1	3
4.	boiler troubles (scale and sludge)	T1	1.12-1.16	BB	1	4
5.	treatment of boiler feed water	T1	1.17-1.18	BB	1	5
6.	Internal treatment (phosphate, colloidal, sodium aluminate and calgon conditioning)	T1	1.18-1.22	BB	1	6
7.	external treatment ,Ion exchange process	T1	1.22-1.27	BB	1	7
8.	zeolite process , desalination of brackish water	W5	1.27-1.30	PPT	1	8
9.	Reverse Osmosis	T1	1.30-1.35	BB	1	9

LEARNING OUTCOME:

At the end of unit , the students will be able to

- Understand the importance of water and its purification.

UNIT II SURFACE CHEMISTRY AND CATALYSIS (9)

10.	Adsorption: Types of adsorption Knowledge about surface chemistry and catalysis (CBS)	T2	32-38	BB PPT	1	10
11.	adsorption of gases on solids – adsorption of solute from solutions	T2	39-45	BB	1	11
12.	adsorption isotherms – Freundlich's adsorption isotherm ,Langmuir's adsorption isotherm	T2	45-48	BB	1	12
13.	contact theory – kinetics of surface reactions, unimolecular reactions	R1	85-89	BB	1	13
14.	Langmuir - applications of adsorption on pollution abatement	R1	90-94	BB	1	14
15.	Catalysis: Catalyst , types of catalysis – criteria – autocatalysis , catalytic poisoning and catalytic promoters	R4	78-83	BB	1	15
16.	acid base catalysis , applications (catalytic convertor	R4	84-90	PPT	1	16
17.	enzyme catalysis	R4	90-94	BB	1	17
18.	Michaelis , Menten equation..	T4	95-98	BB	1	18

LEARNING OUTCOME:

At the end of unit, the students will be able to

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- Understand the importance of surface chemistry and catalysis

UNIT – III
ALLOYS AND PHASE RULE
(9)

19.	Alloys: Introduction- Definition- properties of alloys corrosion of metal alloys and importance in engineering applications in all disciplines.(CBS)	R4	45-48	PPT	1	19
20.	significance of alloying, functions and effect of alloying elements	R4	49-52	BB	1	20
21.	Nichrome and stainless steel (18/8)	T1	3.10-3.12	BB	1	21
22.	heat treatment of steel Basic concepts of alloys and phase rule. (CBS)	T1	3.12-3.15	PPT	1	22
23.	Phase rule: Introduction, definition of terms with examples	T1	4.1-4.2	BB	1	23
24.	one component system -water system	T1	4.2-4.8	BB	1	24
25.	reduced phase rule	T1	4.8-4.10	BB	1	25
26.	thermal analysis and cooling curves	T1	4.10-4.12	BB	1	26
27.	two component systems - lead-silver system- Pattinson process.	T1	4.12-4.14	PPT	1	27

LEARNING OUTCOME:
At the end of unit, the students will be able to

- Strong foundational knowledge in phase rule and Alloys.

UNIT-IV
FUELS AND COMBUSTION
(9)

28.	Fuels: Introduction: Classification of fuels; Coal and coke	R4	125-127	BB	1	28
29.	analysis of coal (proximate and ultimate)	R4	128-130	BB	1	29
30.	carbonization - manufacture of metallurgical coke (Otto Hoffmann method)	W1	85-89	BB	1	30
31.	petroleum - manufacture of synthetic petrol (Bergius process)	T2	89-91	BB	1	31
32.	knocking , octane number , diesel oil , cetane number	R4	145-148	BB	1	32
33.	natural gas - compressed natural gas (CNG) , liquefied petroleum gases (LPG) power alcohol and biodiesel	R3	75-81	PPT	1	33
34.	Combustion of fuels: Introduction , calorific value , higher and lower calorific values	R3	81-85	BB	1	34
35.	ignition temperature , spontaneous explosive range Knowledge about basic in fuels and combustion.(CBS)	R3	86-89	PPT	1	35

36.	flue gas analysis (ORSAT Method)	R3/T2	90-93 189-192	BB	1	36
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LEARNING OUTCOME:

At the end of unit , the students will be able to

- Understand the importance of fuels and combustion.

UNIT-V ENERGY SOURCES AND STORAGE DEVICES

(9)

37.	Nuclear fission - controlled nuclear fission	R3	56-57	BB	1	37
38.	nuclear fusion - differences between nuclear fission and fusion	T3	221-225	BB	1	38
39.	nuclear chain reactions - nuclear energy	T3	225-229	BB	1	39
40.	light water nuclear power plant - breeder reactor	T3	230-235	PPT	1	40
41.	solar energy conversion - solar cells, wind energy	W4	125-128	BB	1	41
42.	Batteries, fuel cells and supercapacitors	R3	128-130	BB	1	42
43.	Types of batteries – primary battery (dry cell)	W3	275-279	PPT	1	43
44.	secondary battery (lead acid battery, lithium-ion-battery) fuel cells	R4	279-282	BB	1	44
45.	H ₂ -O ₂ fuel cell Basic concept of energy sources and storage devices.(CBS)	R4	283-288	BB PPT	1	45

LEARNING OUTCOME:

At the end of unit , the students will be able to

- Comprehend and apply energy sources and applications of energy storage devices.

COURSE OUTCOMES:

At the end of the course, the students will be able:

C104.1: Summarize the water related problems in boilers and their treatment techniques.

C104.2: Discuss the applications of adsorption in the field of water and air pollution abatement.

C104.3: Discuss the types of catalysis and the mechanism of enzyme catalysis

C104.4: Associate phase rule in the alloying and the behaviour of one component and two components Systems using phase diagram

C104.5: Explain various types of fuels, their manufacturing processes and calculation of calorific theoretically

C104.6: Summarize the principles and generation of energy in batteries, nuclear reactors, solar cells, Wind mills and fuel cells

CONTENT BEYOND THE SYLLABUS

- Importance of water and its purification methods
- Knowledge about surface chemistry and catalysis
- Basic concepts of alloys and phase rule.
- corrosion of metal alloys and importance in engineering applications in all disciplines.
- Knowledge about basic in fuels and combustion.
- Basic concept of energy sources and storage devices .

Dr. S.THILAGAVATHI M.E., Ph.D.,
PRINCIPAL

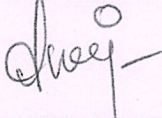
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CONTINUES INTERNAL ASSESSMENT DETAILS

ASSEMET NUMBER	I	II
UNIT	1 st , 2 nd , 3 rd (Half)	3 rd (Half), 4 th & 5 th Units

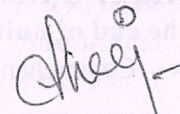
ASSIGNMET NUMBER	UNIT	ASSIGNMENT QUESTIONS/TOPIC	DEADLINE
1	I	1. estimation of hardness of water by EDTA 2. Revers osmosis	21/10/2018
2	II	1. Freundlich's adsorption isotherm 2. Catalyst – types of catalysis , criteria , autocatalysis , catalytic poisoning and catalytic promoters .	07/11/2018
3	III	1. stainless steel and explain its types. 2. condensed phase rule, two-component system	09/12/2019
4	III	1. Proximate, ultimate analysis of coal 2. Problems based on fuels.	02/01/2019
5	IV	1. Problems based energy sources. 2. Importance of battery	22/01/2019

PREPARED BY



Mrs. T. ANNALAKSHMI, AP/CHEMISTRY

VERIFIED BY



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KAIKKURICHI
PUDUKKOTTAI - 622 303.

APPROVED BY

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PRINCIPAL

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COLLEGE FOR WOMEN
KAIKKURICHI - 622 303.
PUDUKKOTTAI DISTRICT

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DEPARTMENT OF SCIENCE AND HUMANITIES

Ref: SBECW/ S&H/ Course committee meeting / EC/ 2018-2019 (ODD)

DATE: 03.09.2018

COURSE COMMITTEE MEETING- CY8151: Engineering chemistry

ACADEMIC YEAR: 2018-2019 (ODD)

PROGRAM : B.E –CSE

DATE OF MEETING : 03.09.2018

REGULATION: 2017

TIME : 10.00AM

SEM : 01

VENUE : S&H Dept. HoD Cabin

Members Present

Table.1 Course committee members

S.No.	Name of the faculty & Designation, Program	Sem/Sec/Program	Signature
1.	Mrs.T.ANNALAKSHMI, HoD/S&H - Course coordinator	I SEM/A/CSE	
2.	Mrs.S.RENUGADEVI AP/Chem	I SEM/B/ECE,CIVIL,EEE	

HOD welcomed all the members present

1. Unit wise syllabus discussed. Nature of qualitative, quantitative, problematic, theoretical concepts etc. have been discussed.

Table.2 Allocation of Period

Number of period per unit	Total number of Peroids
09	45

2. Vision and mission of the college, department discussed. POs, PEOs, PSOs discussed.
3. Course outcomes finalized for each units.

Table.3 Course Outcomes

CO	Course Outcomes	POs	PSOs
C104.1	Summarize the water related problems in boilers and their treatment techniques.	1,2,3,4,7,10,12	1
C104.2	Discuss the applications of adsorption in the field of water and air pollution abatement.	1,2,3,4,7,10,12	1
C104.3	Discuss the types of catalysis and the mechanism of enzyme catalysis	1,2,3,4,7,10,12	1
C104.4	Associate phase rule in the alloying and the behaviour of one component and two component systems using phase diagram	1,2,3,4,7,10,12	1
C104.5	Explain various types of fuels, their manufacturing processes and calculation of calorific theoretically	1,2,3,4,7,10,12	1
C104.6	Summarize the principles and generation of energy in batteries, nuclear reactors, solar cells, wind mills and fuel cells	1,2,3,4,7,10,12	1

4. Mapping of COs with POs and PSOs is done with suitable correlation levels (1 for low, 2 for medium, 3 for high, "-" for no correlation, before content beyond syllabus)

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Table.4 Mapping of COs, C, PSOs with POs- before CBS.

Course	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2
C104.1	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.2	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.3	3	3	1	1	-	-	1	-	-	1	-	1	1	-
C104.4	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.5	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.6	3	2	1	1	-	-	1	-	-	1	-	1	1	-

5. Identification of content beyond syllabus- curricular gaps are identified considering industry needs, employers feedback, alumni feedback, government policy on industrialization, new investments by private/ public sectors, societal needs and level of correlation of COs with POs and PSOs. Accordingly the details of CBS added and its correlation is given below.

Table.5 Identification of content beyond syllabus

Content beyond syllabus added	POs strengthened/Vacant filled	CO/Unit
corrosion of metal alloys and importance in engineering applications in all disciplines..	PO4(1) strengthened	C104.2/ II

6. Mapping of COs with POs, PSOs- after CBS.

Table.6 Mapping of COs, C, PSOs with POs- after CBS.

Course	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2
C104.1	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.2	3	2	1	*2	-	-	1	-	-	1	-	1	1	-
C104.3	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.4	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.5	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104.6	3	2	1	1	-	-	1	-	-	1	-	1	1	-
C104	3	2	1	1	-	-	1	-	-	1	-	1	1	-


7. Content beyond syllabus is thus identified based on the above. Plan for handling of CBS by internal/external resource person/ industrial visits are decided. This will be included in the class log book.
8. Lecture schedule should be prepared unit wise, as in the syllabus. Number of periods per unit and total number of periods planned should not be less than, periods allotted in the syllabus of Anna University.
9. Plan for additional Periods for IA tests, CBS, NPTEL delivery, Seminar, Quiz etc are to be incorporated in the lecture schedule. These periods are added exclusive of number of periods prescribed in the syllabus.
10. Plan for at least three assignments (with level of correlation), seminar topic, quiz questions discussed.
11. Bright students and slow learners are to be identified, immediately after IA test - I. such students may be counselled suitably and the evidence for counselling to be recorded in the attendance cum assessment record.

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(Sign of students with date and time of counselling, to be strictly recorded and to be attached in the course file).

12. For those students secured less than 60% in the IA Test, Makeup test should be conducted. Correspondingly root cause analysis for reasons of failure, corrective and preventive action, and follow up action taken should be filed properly.
13. Contents of course file to be reviewed periodically.
14. Lecture schedule, assignment questions, tutorial questions, course materials, AU questions (at least 5) should be supplied within one week after the commencement of classes.
15. Course material should be uploaded in the college website for student's reference.
16. Discrepancy in question paper, if any to be informed to the controller of examinations through web portal entry, after getting approval from the HoD & the Principal. Critically asked questions, if any to be discussed with the students of the next batch.
17. Immediately after the publication of the results, analysis are to be carried out and follow up action to be taken for the failures.
18. IA test question papers should be set as per the norms of the college, incorporating marks for learning outcomes and course outcomes. Common question papers should be set.
19. Certificate course/guest lectures may be planned inviting experts from industry/higher learning institutions.
20. After IA test, an objective type tests may be conducted (3 times in a semester-30 minutes duration-maximum 10 questions). Questions asked in GATE, TANCET, IES or any other Competitive examination can be taken as a reference. This is to facilitate the bright students to prepare for higher level of thinking and to enhance placement and higher studies opportunities.
21. IA test papers, assignment papers or any other papers submitted by the students, should be returned to the students within 5 days after correction. Sample paper should be suitably filed.
22. Long absentees of students if any to be informed to the parents through class coordinator, if such students attendance less than 75%.


COURSE COORDINATOR


HoD/S&H
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PRINCIPAL
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COLLEGE FOR WOMEN
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Kaikkurichi, Pudukkottai, Tamil Nadu – 622 303, India

DEPARTMENT OF SCIENCE AND HUMANITIES

Identification of Curricular Gap & Content Beyond Syllabus(CBS)

Name of the Faculty : T.ANNALAKSHMI, AP/CHEMISTRY

Course Code & Name: CY8151 ENGINEERING CHEMISTRY

Degree & Program : B.E. /CSE Semester: I Academic Year: 2018 -2019 /ODD

I. Mapping of Course Outcomes with POs & PSOs. (before CBS)

Table.1 Mapping of COs, C, PSOs with POs - before CBS.

CO/PSO	PO 1	PO 2	PO 3	PO 4	PO 5	PO 6	PO 7	PO 8	PO 9	PO 10	PO 11	PO 12	PSO 1	PSO 2	PSO 3
C104.1	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.2	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.3	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.4	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.5	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.6	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-

II. Identification of content beyond syllabus.

Table.2 Identification of content beyond syllabus

Details of Content Beyond Syllabus(CBS) added	POs strengthened/ vacant filled	CO/Unit
Corrosion of Metal Alloys and importance in engineering applications in all disciplines.	PO4(1) strengthened	C104.2/ II

III. Mapping of Course Outcomes with POs & PSOs. (After CBS)

Table.3 Mapping of COs, C, PSOs with POs- after CBS.

CO/PSO	PO 1	PO 2	PO 3	PO 4	PO 5	PO 6	PO 7	PO 8	PO 9	PO 10	PO 11	PO 12	PSO 1	PSO 2	PSO 3
C104.1	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.2	3	2	1	*2	-	-	1	-	-	1	-	1	1	-	-
C104.3	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.4	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.5	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.6	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-

Signature of the Faculty

Dr. S.THILAGAVATHI M.E., Ph.D.,

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TOPICS BEYOND THE SYLLABUS

CORROSION OF THE METAL ALLOYS

REASON BEYOND THE SELECTION

To acquire knowledge of corrosion of metal alloys and importance in engineering applications and metal alloys

REFERENCE LINK

https://www.lkouniv.ac.in/site/writereaddata/siteContent/202003251903229434ranvijay_Corrosion_Science.pdf

Course	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO
C101	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C102	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C103	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C104	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C105	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C106	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C107	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C108	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C109	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C110	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1

Column	POs (strengthened/ vacant filled)	Details of Content Beyond Syllabus (CBS) added
C104A II	PO4 (I) strengthened	Corrosion of Metal Alloys and importance in engineering applications in all disciplines

Course	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO	PO
C101	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C102	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C103	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C104	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C105	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C106	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C107	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C108	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C109	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1
C110	1	1	1	1	1	1	1	1	1	1	1	1	1	1	1


Dr. S. THILAGAVATHI M.E., Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
Kaikkurchi - 622 303, Pudukkottai Dt.



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DEPARTMENT OF SCIENCE AND HUMANITIES

Assignment Question Paper

Name of the Student: *Gayathri.S*

AU Register Number: *912618104006*

Assignment – 03			Date of Issue:	09.12.2018	Marks	30
Course code	CY8151	Course Title	Engineering chemistry			
Year	I	Semester/Section	I / A	Date of Submission:	12.12.2018	

Q.No	Questions	CO
1	What is stainless steel and explain its types?	C104.3
2	What is condensed phase rule? What is the number of degrees of freedom at eutectic point for a two-component system?	C104.3
3	What are the purposes of alloy making? Illustrate with suitable examples.	C104.3

Name and Signature of the Faculty Incharge

[Signature]
[T. ANNALAKSHMI]

[Signature]
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KAIKKURICHI
PUDUKKOTTAI - 622 303.

[Signature]
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PRINCIPAL
SRI BHARATHI ENGINEERING
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DEPARTMENT OF SCIENCE AND HUMANITIES

Assignment Answer Sheet

Name of the Student: *Crayalathi - S*

AU Register Number: *912618104006*

Assignment -03			Date of Issue:	09.12.2018	Marks	30
Course code	CY8151	Course Title	Engineering chemistry			
Year	I	Semester/Section	I / A	Date of Submission:	12.12.2018	

Q.No	Questions	CO
1	What is stainless steel and explain its types?	C104.3
2	What is condensed phase rule? What is the number of degrees of freedom at eutectic point for a two-component system?	C104.3
3	What are the purposes of alloy making? Illustrate with suitable examples.	C104.3

Mark Allocation

Rubrics	Marks Allocated	Marks obtained
Content Quality	6	5
Presentation Quality	2	2
Timely submission	2	2
Total marks	10	9

Name and Signature of the Faculty Incharge

[T. ANNALAKSHMI]

[Signature]

Dr. S. THILAGAVATHI M.E., Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
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Kaikkurichi - 622 303, Pudukkottai Dt.

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IQAC Academic Audit Form

ACADEMIC YEAR: 2018-2019 ODD SEMESTER

Name of Department : CSE Year / Sem / Sec : I / II / A No. of Students Registered : 31

Details of Examination : CT -1 / CT -2 / CT -3 / Model Test

S.No.	Course Code	List of Reg.No Verified	Course Log Book Verified (Y/N)	Course File Verified (Y/N)	No of students passed	No of Absentees	No of Failures	Pass %	Remarks
1	HS8151	912618104006	Y	Y	06	02	23	20.68	-
2	MA8151	912618104001	Y	Y	11	02	18	37.93	-
3	PH8151	912618104023	Y	Y	17	02	12	58.62	-
4	CY8151	912618104015	Y	Y	27	02	02	93.10	-
5	UTE8151	912618104011	Y	Y	22	02	07	75.80	-
6	UE8152	912618104029	Y	Y	13	03	15	46.42	-

Verified by

External Member Name and Signature:

G. Sugapriya

Internal Member Name and Signature:

[R SARATHI] R. SATHI

Overall Remarks:

Dr. S. THILAGAVATHI M.E., Ph.D.,
PRINCIPAL

SRI BHARATHI ENGINEERING
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HOD / S&H
HOD / S&H

IQAC Coordinator

Principal

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DEPARTMENT OF SCIENCE AND HUMANITIES

SUBJECT CODE & TITLE : CY8151-ENGINEERING CHEMISTRY

YEAR/SEM/SEC/BRANCH : I/I/A/CSE

STUDENT FEEDBACK ON FACULTY

S.NO.	DESCRIPTION	SCORED OUT OF 4	SCORED OUT OF 100
1.	The syllabus of the courses that you have studied in relation to the course competencies expected.	3.7	92
2.	The course curriculum framed by university is relevant to the programme and year of study.	3.5	87
3.	The allocation of the hours and credit to the course.	3.6	90
4.	The electives offered in syllabus are in relation to the technological advancement.	3.7	92
5.	Size of syllabus in terms of the load on the student.	3.5	87
6.	The courses in terms of extra learning or self-learning considering the design of the courses.	3.8	95
7.	The evaluation scheme designed for each of the course.	3.7	92
8.	The co-curricular activities conducted to bridge the gap between Academic and Industrial needs.	3.8	95
OVERALL SCORE		3.6	91


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PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
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PRINCIPAL
SRI BHARATHI ENGINEERING
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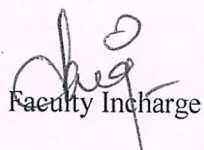
REPORT SHEET

S.NO	REG.NO	NAME	Q1	Q2	Q3	Q4	Q5	Q6	Q7	Q8
1.	912618104001	ABIRAMI.S	4	3	4	3	4	3	4	4
2.	912618104002	AKILA.P	3	4	4	4	4	3	4	4
3.	912618104003	BUVANESHWARI.S	3	4	4	3	4	3	4	4
4.	912618104004	EVANJELIN.S	4	4	4	4	4	4	3	4
5.	912618104005	FEFINA.I	3	3	4	4	3	4	4	4
6.	912618104006	GAYATHRI.S	3	4	4	3	4	4	4	4
7.	912618104007	GOWSALYA.A	4	4	4	4	4	3	4	4
8.	912618104008	ISHWARYA.S	4	3	4	4	3	3	3	4
9.	912618104009	ISWARYA.C	3	4	4	3	4	4	4	4
10.	912618104011	JAYANTHINI.T	4	3	4	3	4	4	3	3
11.	912618104012	KAVIYASELVI.K	3	4	3	3	3	4	3	4
12.	912618104013	KOWSALYA.S	4	4	3	4	4	4	4	4
13.	912618104014	LAKSHMI.N	4	4	3	4	4	3	4	4
14.	912618104015	LAKSHMI PRABHA.M	4	3	4	4	4	4	4	3
15.	912618104016	LAVANYA.S	3	3	4	3	4	4	4	4
16.	912618104017	MALA.S	3	4	4	4	4	4	4	3
17.	912618104018	MAMTHA.G	4	3	4	3	4	4	4	4
18.	912618104019	NIRANJANADEVIC	4	4	3	3	4	3	4	4
19.	912618104020	NIVEDHA.M	3	4	4	4	4	3	3	3
20.	912618104021	NIVETHA.G	4	4	3	3	3	4	4	4
21.	912618104022	PRIYADHARSHINI.R	3	4	4	4	4	4	4	4


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PRINCIPAL
SRI BHARATHI ENGINEERING
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22.	912618104023	PRIYAVATHAN.I.A	4	3	4	3	3	3	4	4
23	912618104024	PUVIYARAS.I.S	3	4	4	4	4	4	4	4
24.	912618104025	RASMI.J	4	4	4	3	4	3	4	4
25.	912618104026	ROSLINA BEGUM.R	3	3	3	4	4	4	4	4
26.	912618104027	SINDHU.V	3	4	4	3	3	4	4	4
27.	912618104028	SIVASANGAVI.A	4	4	4	4	4	4	4	4
28.	912618104029	SURUTHIKA.S	4	3	4	3	3	4	4	4
29.	912618104030	SURYA.A	3	4	4	4	3	4	4	3
30.	912618104031	SUSHMEENA K	4	4	3	3	4	4	4	4
AVERAGE			3.5	3.6	3.7	3.5	3.7	3.6	3.8	3.8
PERCENTAGE			88	91	94	87	93	91	95	95

EXCELLENT	VERY GOOD	GOOD	AVERAGE	POOR
4	3	2	1	0


Faculty Incharge


Dr. S.THILAGAVATHI M.E., Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
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Circular

Date: 22-10-2018

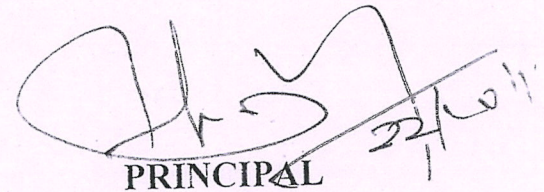
The second cycle test will be conducted on 31.10.2018, 01.11.2018, 02.11.2018 & 03.11.2018 for the I semester (I year) students.

The following instructions are to be followed by the faculty members.

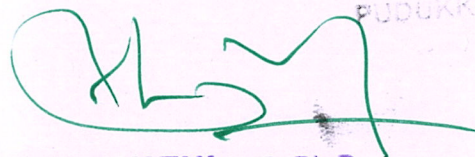
- Total marks for which the question paper to be set will be for 50 marks.
- It is the responsibility of the question paper setter to take the Xerox copies of the required number of question papers with the help of Mr. Pandi. S & Ms. Anusha. G and it should be handed over to the Exam Coordinator Mr. J. Sathyaraj A.P/ EEE on or before 29.10.2018.
- The Exam Coordinators (exam cell) are requested to make necessary arrangements (hall arrangements, invigilation duty etc.,) for conducting the test.
- Faculty members are requested to handover the valued answer scripts to the students on or before 08.11.2018 and the class in-charges are requested to send the consolidated mark sheet along with the attendance percentage to the parents on or before 09.11.2018.

Cc:

- All faculty
- Exam cell
- Office file


PRINCIPAL 22/10/18

PRINCIPAL
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KAIKKURICHI, PUDUKKOTTAI DISTRICT



Dr. **S.THILAGAVATHI M.E., Ph.D.**
PRINCIPAL
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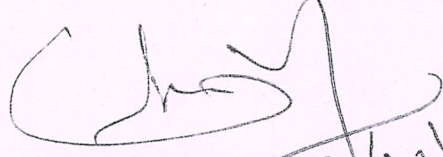
**SRI BHARATHI ENGINEERING COLLEGE FOR WOMEN
KAIKKURICHI, PUDUKKOTTAI – 622 303.**

Circular

Date: 22-10-2018

The second cycle test will be conducted on 31.10.2018, 01.11.2018, 02.11.2018 & 03.11.2018 for the I semester (I year) B.E/ B.Tech students for 50 marks as per the time table given below. Students are directed to prepare well and score good marks.

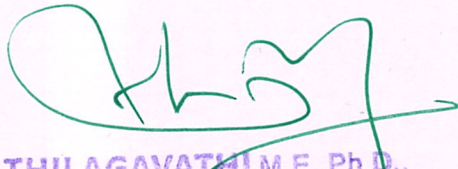
Date	(10.00 am -11.30 am)	2.30 pm -4.00 pm
31.10.2018	CY8151 Engineering Chemistry	GE8151 Problem Solving and Python Programming
01.11.2018	PH8151 Engineering Physics	HS8151 Communicative English
02.11.2018	COACHING	MA8151 Engineering Mathematics – I
03.11.2018	GE8152 Engineering Graphics (9.30 pm to 12.30pm)	


PRINCIPAL 22/10/18

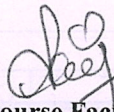
Cc:

- All I year B.E / B.Tech Classes
- All faculty
- Exam cell
- Notice Board
- Office file

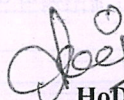
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KAIKKURICHI - 622 303.
PUDUKKOTTAI DISTRICT**


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11a	(i) Discuss the heat treatment of Steel in details. (ii) Give the composition and uses of the following alloys. 1) Nichrome 2) Stainless steel	CO104.3	K2
OR			
11b	(i) Discuss the phase diagram of lead-Silver system and explain briefly write about Pattinson's process. (ii) Discuss the phase diagram of water system and explain the characteristics	CO104.3	K2


Course Faculty

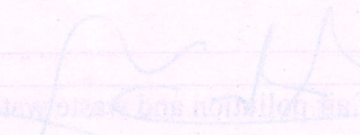
(Name / Sign / Date)


HoD

(Name / Sign / Date)

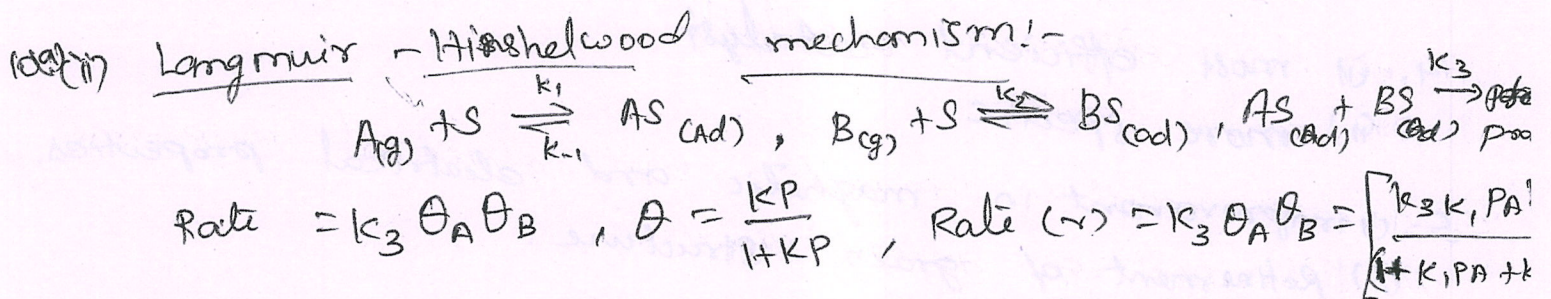
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Mechanism: Step-I - Adsorption of reactant molecules.
 Step-II - Formation of Activated complex.
 Step-III - Decomposition of Activated complex.
 Step-IV - Desorption of products.

- * finely divided state of catalyst is more efficient.
- * Enhanced activity of a rough surfaced catalyst.
- * Action of promoters, catalytic poisoning.
- * specific action of the catalyst.



(i) Both the molecules have low adsorption.

$$1 \gg K_1 P_A, K_2 P_B \therefore \text{Rate (r)} = k_3 K_1 K_2 P_A P_B$$

(ii) One molecule have low adsorption.

$$K_1 P_A \gg 1 \gg K_2 P_B, \text{Rate (r)} = \frac{k_3 K_1 K_2 P_A P_B}{(1 + K_1 P_A)^2}$$

(iii) one molecule have very high adsorption.

$$K_1 P_A \gg 1, K_2 P_B, \text{Rate (r)} = \frac{k_3 K_1 P_A K_2 P_B}{(K_1 P_A)^2}$$

(10.10.11) Air pollution and waste water management:

1. Granular activated carbon (GAC)
2. Powdered Activated carbon (PAC)

PAC 1. down flow fixed-bed carbon contactors.
 2. upflow fixed-bed carbon contactors.

PAC powdered activated carbon (PAC) is added directly into the effluent coming out from the various biological treatment processes

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11. (b) (i) Lead-silver system

$F = C - P + 1$ at constant pressure, the vapour phase is ignored and the condensed phase rule is used.

curve OA Solid Ag \rightleftharpoons melt. $F = 2 - 2 + 1$ $F = 1$

curve BO Solid Pb \rightleftharpoons melt. $F = 2 - 2 + 1$ $F = 1$

point "O" Solid Pb + Solid Ag \rightleftharpoons melt. $F = 2 - 3 + 1$ $F = 0$.

Areas : (molten Pb + Ag) $F = 2 - 1 + 1$ $F = 2$

11. (b) (ii) Water system :-

Solid \rightleftharpoons liquid

liquid \rightleftharpoons vapour

Solid \rightleftharpoons vapour.

curve OA : water \rightleftharpoons water vapour.

$F = C - P + 2$, $F = 1 - 2 + 2$, $F = 1$

curve OB : Ice \rightleftharpoons vapour.

$F = 1 - 2 + 2$, $F = 1$

curve OC : Ice \rightleftharpoons water

$F = 1 - 2 + 2$, $F = 1$

point 'O' Ice \rightleftharpoons water \rightleftharpoons vapour.

$F = C - P + 2$, $F = 1 - 3 + 2$, $F = 0$

curve OB'

Super-cool water \rightleftharpoons vapour.

Areas AOB, BOC, AOC. represents water vapour, ice

$F = C - P + 2$, $F = 2 - 1 + 2$ $F = 3$.


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Cycle Test Answer Book

Name	NIVETHA.M			Year/ Semester/Section	I/I/A			
Batch No.	912618104020	Date/Session	31.10.2018	Department:	CSE			
Course code	LY8151	Course Title	Engineering Chemistry					
Internal Assessment Test	IAT 1	<input type="checkbox"/>	IAT 2	<input checked="" type="checkbox"/>	IAT 3	<input type="checkbox"/>	Model	<input type="checkbox"/>
Name and Signature of the Invigilator with date	S. Devi 31/10/18 [S.RENVUJADEVI]							

Instruction to the Student: Put tick mark to the question attended in the column against question.								
Part A			Part B / Part C				Total Marks	
Q. No.	✓	Marks	Q. NO.	✓	a	b		Total Marks
					Marks	Marks		
1	✓	2	11	✓	13		13	
2	✓	2	12			✓ 12	12	
3	✓	2	13					
4	✓	0	14					
5	✓	2	15					
6	✓	0	16					
7	✓	2	Total				25	
8	✓	2	37 Grand Total				[S. ANNAPPAK] SHML 11/11/2018 Name and Signature of the Examiner with date	
9	✓	0						
10	✓	0						
Total		12						

To be filled by the examiner							
Course Outcomes	1	2	3	4	5	6	Total
Marks allotted		24	26				50
Marks Obtained		19	18				37
IQAC Audit - Remarks							[P. SUBHA]
							Name and Signature of the IQAC member

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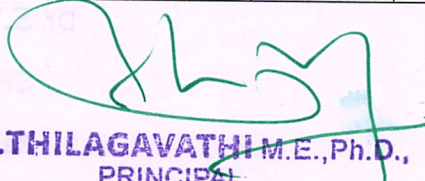
ACADEMIC YEAR 2018-2022-- ODD SEMESTER
STUDENTS MARK STATEMENT –CO BASED

SECTION -A

CYCLE TEST-II-(NOVEMBER-2018)

PROGRAM : B.E / CSE
YEAR/SEM : I/I
SUBJECT CODE & TITLE : CY8151-ENGINEERING CHEMISTRY
DATE : 31.10.2018

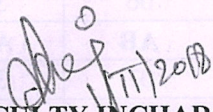
S.NO	REG.NO	NAME	CO3 (10)	CO4 (20)	CO5 (10)	CO6 (10)	MARKS (50)	MARKS (100)
1	912618104001	ABIRAM.I.S	06	10	08	07	31	62
2	912618104002	AKILA.P	05	12	09	07	33	66
3	912618104003	BUVANESHWARI.S	7	15	09	05	36	72
4	912618104004	EVANJELIN.S	08	13	06	03	30	60
5	912618104005	FEFINA.I	04	06	05	01	16	32
6	912618104006	GAYATHRI.S	09	18	08	09	44	88
7	912618104007	GOWSALYA.A	05	08	07	06	26	52
8	912618104008	ISHWARYA.S	09	16	06	05	36	72
9	912618104009	ISWARYA.C	08	15	07	06	36	72
10	912618104011	JAYANTHINI.T	AB	AB	AB	AB	AB	AB
11	912618104012	KAVIYASELVI.K	07	14	08	07	36	72
12	912618104013	KOWSALYA.S	06	13	09	04	32	64
13	912618104014	LAKSHMI.N	09	18	08	07	42	84
14	912618104015	LAKSHMI PRABHA.M	08	14	07	09	38	76
15	912618104016	LAVANYA.S	08	18	09	06	41	82
16	912618104017	MALA.S	06	14	06	04	30	60
17	912618104018	MAMTHA.G	AB	AB	AB	AB	AB	AB
18	912618104019	NIRANJANADEVIC	08	15	08	08	39	78
19	912618104020	NIVEDHA.M	08	14	08	07	37	74
20	912618104021	NIVETHA.G	06	14	07	04	31	62
21	912618104022	PRIYADHARSHINI.R	06	19	08	08	41	82
22	912618104023	PRIYAVATHANI.A	06	16	06	05	36	72
23	912618104024	PUVIYARASI.S	04	06	05	04	19	38
24	912618104025	RASMI.J	07	13	06	04	30	60
25	912618104026	ROSLINA BEGUM.R	09	18	10	09	46	92
26	912618104027	SINDHU.V	09	18	09	09	45	90
27	912618104028	SIVASANGAVI.A	07	19	07	07	40	80
28	912618104029	SURUTHIKA.S	06	14	06	05	31	62
29	912618104030	SURYA.A	07	19	08	09	43	86
30	912618104011	SUSHMEENA K	06	10	04	07	27	54

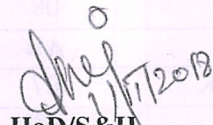

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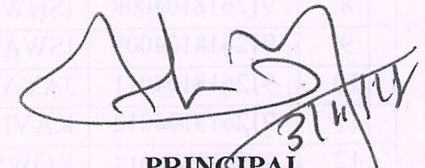
MARK RANGE:


<20	20-30	31-40	41-50	51-60	61-70	71-80	81-90	91-100
-	-	02	-	05	05	09	06	01

Total Number of Students Present	28
Total Number of Students Absent	02
Total Number of Candidates Pass	26
Total Number of Candidates Fail	02
Percentage of Pass	92.85%


11/11/2018
FACULTY INCHARGE


11/11/2018
HOD / S&H
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PUDUKKOTTAI - 622 303.


31/11/18
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Kaikkurichi, Pudukkottai, Tamil Nadu – 622 303, India


DEPARTMENT OF SCIENCE AND HUMANITIES

ROOT CAUSE ANALYSIS

Name of the Faculty : Mrs.T.Annalakshmi Course Code & Name : CY8151-Engineering Chemistry
Degree & Program : B.E & CSE Semester : I
Cycle Test : II Month & Year : November / 2018
Target : 100 % Achieved : 92.85%

S.NO	REG NO	NAME OF THE STUDENT	CAUSES FOR FAILURE	CORRECTIVE ACTION TAKEN
1.	912618104005	FEFINA.I	confused in derivation part	during office hours to ask questions about lectures to the faculty
2.	912618104007	GOWSALYA.A	Not prepare well	Home test given
3.	912618104011	JAYANTHINI.T	ABSENT	counseling given to attend test with well preparation
4.	912618104018	MAMTHA.G	ABSENT	counseling given to attend test with well preparation
5.	912618104024	PUVIYARASI.S	Confused in equation part	during office hours test conducted
6.	912618104031	SUSHMEENA K	Not prepare well	Home test given.


Signature of the Faculty Member


Signature of the HoD/S&H


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Circular

Date: 05.11.2018

Retest for second cycle test will be conducted from **09.11.2018** to **15.11.2018** for the I semester (I year) students.


The following instructions are to be followed by the faculty members.

The following instructions are to be followed by the faculty members.

- Total marks for which the question paper to be set will be for 50 marks.
(PART A 5X2=10, PART B 2X13=26 & PART C 1X14=14)
- It is the responsibility of the **question paper** setter to take the Xerox copies of the required number of question papers.
- Concerned Faculty members are requested to conduct the examination as per the scheduled and handover the valued answer scripts to the students on or before **16.11.2018**.

Cc:

- All faculty
- Exam cell
- Office file


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Circular

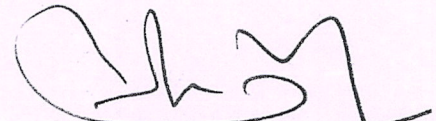
Date: 06.11.2018

Retest for second cycle test will be conducted from **09.11.2018** to **15.11.2018** for the I semester (I year) B.E students for **50 marks** as per the time table given below. Students are directed to prepare well and score good marks.

Date	04.00 pm -05.30 pm
09.11.2018	MA8151-Engineering Mathematics-I
10.11.2018	PH8151-Engineering Physics
12.11.2018	HS8151-Communicative English
13.11.2018	GE8152-Engineering Graphics
14.11.2018	CY8151-Engineering Chemistry
15.11.2018	GE8151-Problem Solving and Python Programming

Cc:

- All I year B.E Classes
- All faculty
- Exam cell
- Notice Board
- Office file


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06/11/18

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ACADEMIC YEAR 2018-2022-- ODD SEMESTER
STUDENTS MARK STATEMENT -CO BASED
SECTION -A

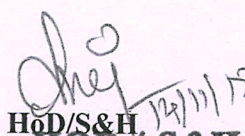
ATTENDANCE SHEET- RETEST FOR CYCLE TEST-II

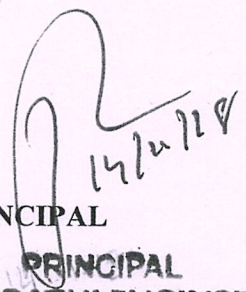
PROGRAM : B.E / CSE
YEAR/SEM : II
SUBJECT CODE & TITLE : CY8151-ENGINEERING CHEMISTRY
DATE : 14.11.2018

S.NO	REG.NO	NAME	SIGNATURE
1	912618104005	FEFINA.I	Fefina.I
2	912618104007	GOWSALYA.A	Gowsalya.A
3	912618104011	JAYANTHINI.T	Jayanthini.T
4	912618104018	MAMTHA.G	Mamtha.G
5	912618104024	PUVIYARASI.S	Puviyarasi.S
6	912618104011	SUSHMEENA K	Sushmeena.K


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FACULTY INCHARGE


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ACADEMIC YEAR 2018-2022-- ODD SEMESTER
STUDENTS MARK STATEMENT –CO BASED
SECTION -A
CYCLE TEST-II (RETEST) (NOVEMBER-2018)

PROGRAM : B.E / CSE
YEAR/SEM : I/I
SUBJECT CODE & TITLE : CY8151-ENGINEERING CHEMISTRY
DATE : 14.11.2018

S. NO	REG.NO	NAME	CO3 (10)	CO4 (20)	CO5 (10)	CO6 (10)	MARKS (50)	MARKS (100)
1	912618104005	FEFINA.I	06	15	05	08	34	68
2	912618104007	GOWSALYA.A	08	15	07	06	36	72
3	912618104011	JAYANTHINI.T	07	14	08	06	35	70
4	912618104018	MAMTHA.G	07	16	08	08	38	76
5	912618104024	PUVIYARASIS	08	18	07	06	39	78
6	912618104011	SUSHMEENA K	07	16	08	08	42	76

MARK RANGE:

<20	20-30	31-40	41-50	51-60	61-70	71-80	81-90	91-100
-	-	-	-	-	02	04	-	-

Total Number of Students Present	06
Total Number of Students Absent	NIL
Total Number of Candidates Pass	06
Total Number of Candidates Fail	NIL
Percentage of Pass	100%

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DEPARTMENT OF SCIENCE AND HUMANITIES

ACADEMIC YEAR 2018-2019 (ODD SEMESTER)

FINAL INTERNAL STUDENTS MARK STATEMENT(Out of 20)

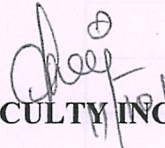
SUBJECT CODE & TITLE: CY8151-ENGINEERING CHEMISTRY

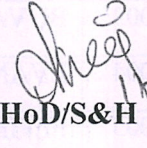
YEAR/SEM/BRANCH: I/I/CSE

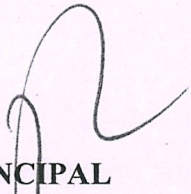
S.NO	REG NO	STUDENT NAME	TOTAL (20)
1.	912618104001	ABIRAMI.S	15
2.	912618104002	AKILA.P	16
3.	912618104003	BUVANESHWARI.S	16
4.	912618104004	EVANJELIN.S	17
5.	912618104005	FEFINA.I	16
6.	912618104006	GAYATHRI.S	17
7.	912618104007	GOWSALYA.A	15
8.	912618104008	ISHWARYA.S	15
9.	912618104009	ISWARYA.C	16
10.	912618104010	JANAKI.V	0
11.	912618104011	JAYANTHINI.T	16
12.	912618104012	KAVIYASELVI.K	17
13.	912618104013	KOWSALYA.S	18
14.	912618104014	LAKSHMI.N	17
15.	912618104015	LAKSHMI PRABHA.M	17
16.	912618104016	LAVANYA.S	12
17.	912618104017	MALA.S	16
18.	912618104018	MAMTHA.G	16


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19.	912618104019	NIRANJANADEVI.C	16
20.	912618104020	NIVEDHA.M	16
21.	912618104021	NIVETHA.G	16
22.	912618104022	PRIYADHARSHINI.R	15
23.	912618104023	PRIYAVATHANI.A	15
24.	912618104024	PUVIYARASI.S	16
25.	912618104025	RASMI.J	18
26.	912618104026	ROSLINA BEGUM.R	18
27.	912618104027	SINDHU.V	16
28.	912618104028	SIVASANGAVI.A	16
29.	912618104029	SURUTHIKA.S	17
30.	912618104030	SURYA.A	15
31.	912618104031	SUSHMEENA K	17

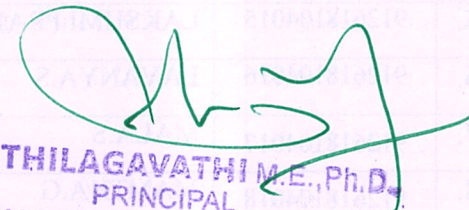

FACULTY INCHARGE

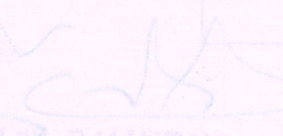

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PUDUKKOTTAI DISTRICT


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SRI BHARATHI ENGINEERING COLLEGE FOR WOMEN

(Approved by AICTE, New Delhi, Affiliated to Anna University, Chennai-25)

Kaikkurichi, Pudukkottai, Tamil Nadu – 622 303, India

DEPARTMENT OF SCIENCE AND HUMANITIES

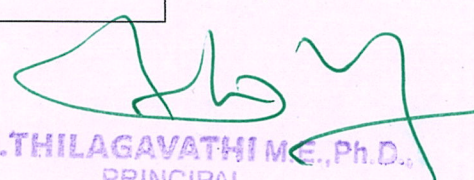
ACADEMIC YEAR 2018 – 2019 (ODD SEMESTER)

ANNA UNIVERSITY RESULT STATEMENT (NOV/DEC- 2018)

SUBJECT CODE & TITLE: CY8151-ENGINEERING CHEMISTRY

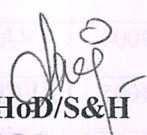
YEAR/SEM/BRANCH: I/I/CSE

S.NO	REG NO	STUDENT NAME	GRADE
1.	912618104001	ABIRAMI.S	B
2.	912618104002	AKILA.P	U
3.	912618104003	BUVANESHWARI.S	B
4.	912618104004	EVANJELIN.S	U
5.	912618104005	FEFINA.I	U
6.	912618104006	GAYATHRI.S	B
7.	912618104007	GOWSALYA.A	U
8.	912618104008	ISHWARYA.S	B
9.	912618104009	ISWARYA.C	B
10.	912618104011	JAYANTHINI.T	B+
11.	912618104012	KAVIYASELVI.K	B
12.	912618104013	KOWSALYA.S	B
13.	912618104014	LAKSHMI.N	B
14.	912618104015	LAKSHMI PRABHA.M	B
15.	912618104016	LAVANYA.S	UA
16.	912618104017	MALA.S	U


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17.	912618104018	MAMTHA.G	B
18.	912618104019	NIRANJANADEVIC	U
19.	912618104020	NIVEDHA.M	U
20.	912618104021	NIVETHA.G	B
21.	912618104022	PRIYADHARSHINI.R	U
22.	912618104023	PRIYAVATHANI.A	B
23.	912618104024	PUVIYARASIS	U
24.	912618104025	RASMI.J	B+
25.	912618104026	ROSLINA BEGUM.R	B+
26.	912618104027	SINDHU.V	U
27.	912618104028	SIVASANGAVI.A	B
28.	912618104029	SURUTHIKA.S	B
29.	912618104030	SURYA.A	U
30.	912618104011	SUSHMEENA K	U


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PRINCIPAL
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SRI BHARATHI ENGINEERING COLLEGE FOR WOMEN, KAIKURUCHI, PUDUKKOTTAI

Department of Science and Humanities

Internal Assessment - Attainment of Course Outcomes (Through Direct Assessment)

ACADEMIC YEAR - 2018 - 19

BATCH

2018 - 2022

COURSE CODE/TITLE

CY8151 (C104)/ENGINEERING CHEMISTRY

COURSE OUTCOME

1 2 3 4 5 6

YEAR/SEM

I/I

TARGET(%)

65 65 65 65 65 65

COURSE COORDINATOR

MRS. T. ANNALAKSHMI

TOTAL STRENGTH

30

ATTAINMENT LEVEL

Level

Range

1

UP TO 60% of the students scored more than target

2

61 - 79% of the students scored more than target

3

80% & ABOVE of the students scored more than target

S.NO	REG NO	NAME OF THE STUDENT	IAT 1 - MARKS ALLOTTED						IAT 2 - MARKS ALLOTTED						IAT 3 - MARKS ALLOTTED						Assignment / Mini Project / Tutorial / Seminar						TOTAL COURSE OUTCOME					
			C1	C2	C3	C4	C5	C6	C1	C2	C3	C4	C5	C6	C1	C2	C3	C4	C5	C6	C1	C2	C3	C4	C5	C6	C1	C2	C3	C4	C5	C6
			60	40							40	60							60	40			10	10					10	60	50	50
1	912618104001	ABIRAMIS	45	30						40	60					60	40			10	10				10	60	50	50	60	60	50	
2	912618104002	AKILA P	45	30						30	45					50	33			8	8				7	45	38	38	45	50	40	
3	912618104003	BUVANESHWARI S	44	30						30	45					50	33			9	9				8	44	39	39	45	50	41	
4	912618104004	EVANJELIN S	48	32						32	47					53	36			7	9				8	48	39	41	47	53	44	
5	912618104005	FEFINAI	47	32						30	45					51	34			8	8				9	47	40	38	45	51	43	
6	912618104006	GAYATHRI S	51	34						35	52					54	36			8	7				9	51	42	42	52	54	45	
7	912618104007	GOWSALYA A	45	30						28	43					51	34			8	8				9	45	38	36	43	51	43	
8	912618104008	ISHWARYA S	43	29						29	44					48	32			9	8				9	43	38	37	44	48	41	
9	912618104009	ISWARYA C	47	32						32	47					53	36			9	9				9	47	41	41	47	53	45	
11	912618104011	JAYANTHINI T	43	28						32	49					52	35			8	8				8	43	36	40	49	52	43	
12	912618104012	KAVIYASELVI K	52	35						34	50					55	36			8	9				8	52	43	43	50	55	44	
13	912618104013	KOWSALYA S	54	36						34	51					56	37			9	9				8	54	45	43	51	56	45.2	
14	912618104014	LAKSHMIN	49	33						35	52					51	34			9	9				8	49	42	44	52	51	42	
15	912618104015	LAKSHMI PRABHA M	48	32						33	50					52	35			8	9				9	48	40	42	50	52	44	
16	912618104016	LAVANYA S	31	20						25	37					37	25			8	9				9	31	28	34	37	37	34	
17	912618104017	MALA S	45	30						32	47					53	35			8	8				8	45	38	40	47	53	43	
18	912618104018	MAMTHA G	50	33						31	46					50	33			8	8				7	50	41	39	46	50	40	
19	912618104019	NIRANJANADEVIC	47	32						33	49					55	37			8	8				8	47	40	41	49	55	45	
20	912618104020	NIVEDHA M	49	32						32	47					50	33			7	7				8	49	39	39	47	50	41	

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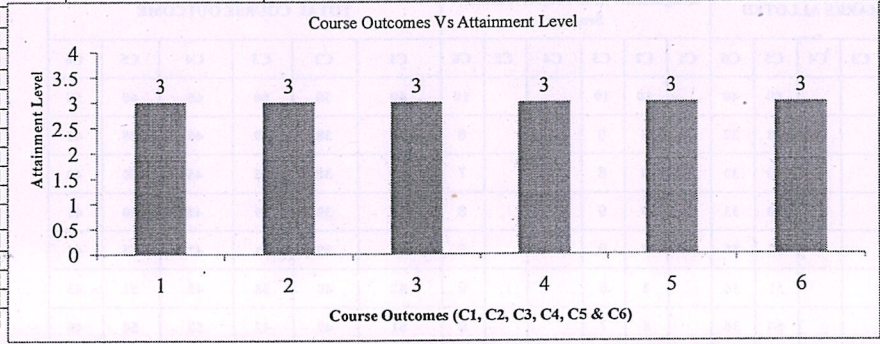
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21	912618104021	NIVETHA G	47	32											33	50							51	34		8	8			9	47	40	41	50	51	43
22	912618104022	PRIYADHARSHINI R	39	26											30	45							52	34		8	9			9	39	34	39	45	52	43
23	912618104023	PRIYAVATHANI A	43	28											29	44							52	35		9	8			9	43	37	37	44	52	44
24	912618104024	PUVIYARASI S	47	32											32	47							53	36		8	8			9	47	40	40	47	53	45
25	912618104025	RASMI J	49	32											37	56							57	38		9	8			8	49	41	45	56	57	46
26	912618104026	ROSLINA BEGUM R.	48	32											36	55							56	37		8	8			8	48	40	44	55	56	45
27	912618104027	SINDHU V	47	32											32	48							51	34		9	9			8	47	41	41	48	51	42
28	912618104028	SIVASANGAVI A	47	32											32	47							50	33		8	9			9	47	40	41	47	50	42
29	912618104029	SURUTHIKA S	51	34											34	51							51	34		8	8			8	51	42	42	51	51	42
30	912618104030	SURYA A	44	29											30	45							49	33		9	8			8	44	38	38	45	49	41
31	912618104031	SUSHMEENA K	51	34											34	52							53	36		9	9			9	51	43	43	52	53	45

CO's Target Value	39.0	33	33	39	39	33
No. of Students scored above CO's Target Value	29	29	30	29	29	30
Percentage of Students scored above Target	97	97	100	97	97	100
CO Attainment	3	3	3	3	3	3
CO attainment Values to plot the Graph	3	3	3	3	3	3



Faculty in charge

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Dr. S. Thilagavathi
Dr. S. THILAGAVATHI M.E., Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
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BHARATHIYAR INSTITUTE OF ENGINEERING FOR WOMEN
DEPARTMENT OF S&H
COURSE OUTCOME ATTAINMENT - UNIVERSITY EXAMINATION
ACADEMIC YEAR : 2018 - 2019 (ODD SEM)

CLASS / SEC: I CSE / A

Batch:2018-2022

SUBJECT :CY8151 (C104) / ENGINEERING CHEMISTRY

CO Attainment Level: 1 - (UPTO 60%) 2- (61%-79%) 3-(80% and Above)

TOTAL STRENGTH : 30

S.NO	Register No	NAME	Univ. Grade	
1	912618104001	ABIRAMI S	B	
2	912618104002	AKILA P	U	
3	912618104003	BUVANESHWARI S	B	
4	912618104004	EVANJELIN S	U	
5	912618104005	FEFINA I	U	
6	912618104006	GAYATHRI S	B	
7	912618104007	GOWSALYA A	U	
8	912618104008	ISHWARYA S	B	
9	912618104009	ISWARYA C	B	
10	912618104011	JAYANTHINI T	B+	
11	912618104012	KAVIYASELVI K	B	
12	912618104013	KOWSALYA S	B	
13	912618104014	LAKSHMI N	B	
14	912618104015	LAKSHMI PRABHA M	B	
15	912618104016	LAVANYA S	UA	
16	912618104017	MALA S	U	
17	912618104018	MAMTHA G	B	
18	912618104019	NIRANJANADEVIC	U	
19	912618104020	NIVEDHA M	U	
20	912618104021	NIVETHA G	B	
21	912618104022	PRIYADHARSHINI R	U	
22	912618104023	PRIYAVATHANI A	B	
23	912618104024	PUVIYARASI S	U	
24	912618104025	RASMI J	B+	
25	912618104026	ROSLINA BEGUM R	B+	
26	912618104027	SINDHU V	U	
27	912618104028	SIVASANGAVI A	B	
28	912618104029	SURUTHIKA S	B	
29	912618104030	SURYA A	U	
30	912618104031	SUSHMEENA K	U	
No. of O Grade			0	0
No. of A+ Grade			0	0
No. of A Grade			0	0
No. of B+ Grade			3	3
No. of B Grade			14	14
No. of U Grade			12	12
No. of UA Grade			1	1
Target for course outcome Attainment			60	30
No of students above the target			17	
CO-Attainment University (%)			56.67	

[Signature]
Faculty

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PUDUKKOTTAI - 622 303.

[Signature]
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Overall Attainment Sheet – COs - POs & PSOs attainment calculation

CO	CO-Attainment Internal (CO-INT) (Avg. Attainment of All section) (%)	CO-Attainment University (Avg. Attainment of All section) (%)	Direct CO Attainment (0.20xCO-INT + 0.80xCO-UNI) (%)	CO Attainment Level
C104.1	96.7	56.67	64.7	2
C104.2	96.7	56.67	64.7	2
C104.3	100.0	56.67	65.3	2
C104.4	96.7	56.67	64.7	2
C104.5	96.7	56.67	64.7	2
C104.6	100.0	56.67	65.3	2

Closure of the Quality Loop:

CO	CO-Target for Academic Year						CO Attainment Gap for (%) 16-17	Action Proposed to Bridge the Gap
	14-15		15-16		16-17			
C104.1	65	79.71	65	69	65	64.7	-	-
C104.2	65	79.71	65	71.17	65	64.7	-	-
C104.3	65	79.71	65	63.15	65	65.3	-	-
C104.4	65	79.71	65	75.11	65	64.7	-	-
C104.5	65	79.71	65	73.57	65	64.7	-	-
C104.6	65	79.71	65	68.44	65	65.3	-	-

Expected CO-PO Level

Course	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
C104.1	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.2	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.3	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.4	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.5	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104.6	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
C104	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-

PO Attainment Level

Course	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
C104.1	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104.2	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104.3	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104.4	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104.5	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104.6	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-
C104	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-

Attainment of POs and PSOs:

Course Code	PO1	PO2	PO3	PO4	PO5	PO6	PO7	PO8	PO9	PO10	PO11	PO12	PSO1	PSO2	PSO3
C104	3	2	1	1	-	-	1	-	-	1	-	1	1	-	-
Attainment	2	1.33	0.67	0.67	-	-	0.67	-	-	0.67	-	0.67	0.67	-	-

Comments by Program Coordinator	1. 2.
Remarks by HoD	

Name and Signature of the Faculty Member

T. Annalakshmi
T. ANNALAKSHMI

[Signature]
Dr. S. THILAGAVATHI M.E., Ph.D.,
PRINCIPAL
SRI BHARATHI ENGINEERING
COLLEGE FOR WOMEN
Kaikkurichi - 622 303, Pudukkottai Dt.

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